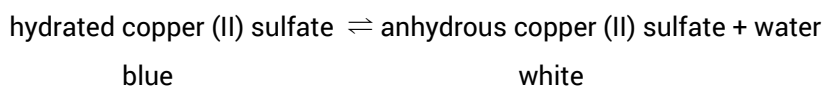


0	1
---	---

The equation below represents the reaction of blue hydrated copper sulfate crystals when they are heated.



0	1
---	---

 .

1

What does \rightleftharpoons tell you about the reaction?

(Reaction is) reversible

[1 mark]

When 10g of blue copper sulfate crystals are heated, 400 joules of energy are taken in.

0	1
---	---

 .

2

What is name of the type of reaction that takes in heat energy?

Endothermic

[1 mark]

0	1
---	---

 .

3

Describe the energy change when an excess of water is added to the anhydrous white copper sulfate.

Energy is transferred to the surroundings or energy is transferred from the anhydrous copper sulfate to the surroundings or heat energy/heat is given off. [1]

[2 marks]

The reaction is exothermic. [1]

(Total 4 marks)

End