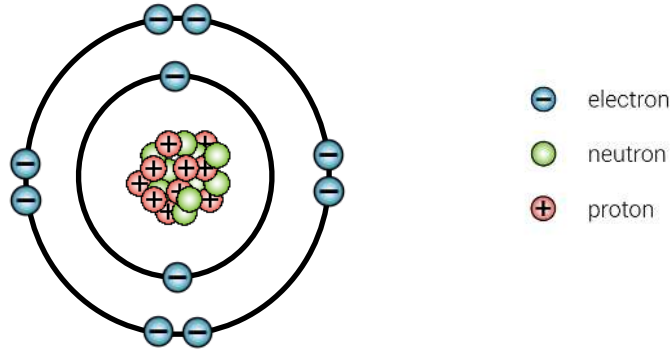


0	1
---	---

The diagram shows a Neon atom.



0	1
---	---

 .

1

Use the correct number to complete each sentence.

10 20 22 40

[2 marks]

The atomic (proton) number of the neon atom shown above is

The mass number of the neon shown above is

0	1
---	---

 .

2

Draw a ring around the correct answer to complete the sentence.

[1 mark]

Neon atoms with different numbers of neutrons are called:

isotopes molecules polymers

0	1
---	---

 .

3

An neon atom with a different number of neutrons has 12 neutrons.

Draw a ring around the symbol which represents this atom.

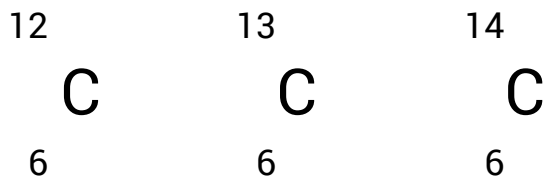
[1 mark]

20	22	24
Ne	Ne	Ne
10	10	10

0

2

There are three isotopes of the element carbon:



0

2

1

Describe, in terms of sub-atomic particles, **one** similarity and **one** difference between atoms of the three isotopes of carbon.

[2 marks]

Similarity

.....

Difference

.....

0

2

2

A sample of charcoal was taken from a newly discovered cave.

The table below show the percentage abundance of carbon in the sample found. Using the data in the table, calculate the relative atomic mass of carbon from this sample:

[2 marks]

Isotope	C ¹²	C ¹³	C ¹⁴
Percentage abundance	80%	15%	5%

.....

.....

.....

Relative Atomic Mass of Carbon:

(Total 8 marks)**End**