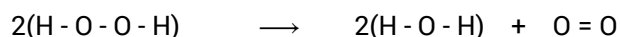


0	1
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Hydrogen peroxide is a chemical that decomposes to produce oxygen and water.

The structural formulae of the chemicals involved are:



**Table 1**

BOND	BOND ENERGY (kJ)
O = O	498
O – O	146
H – O	464

0	1	.	1
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Use the structural formulae above and information from Table1 to calculate the energy required to break the bonds in the reactant.

**[2 marks]**

.....

.....

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0	1	.	2
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Calculate the energy released when new bonds are formed in the products.

**[2 marks]**

.....

.....

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0	1	.	3
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Is the reaction endothermic or exothermic?

**[2 marks]**

.....

Explain why.

.....

.....

.....

.....

.....

**(Total 6 marks)**

**End**